

# Luminescence of BaBrI and SrBrI single crystals doped with Eu<sup>2+</sup>

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## ABSTRACT

The crystal growth procedure and luminescence properties of pure and Eu<sup>2+</sup>-doped BaBrI and SrBrI crystals are reported. Emission and excitation spectra were recorded under ultraviolet and vacuum ultraviolet excitations. The energy of the first Eu<sup>2+</sup> 4f-5d transition and SrBrI band gap are obtained. The electronic structure calculations were performed within GW approximation as implemented in the Vienna Ab Initio Simulation Package. The energy between lowest Eu<sup>2+</sup> 5d state and the bottom of conduction band are found based on luminescence quenching parameters. The vacuum referred binding energy diagram of lanthanide levels was constructed using the chemical shift model.

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## 1. Introduction

For many applications that use inorganic scintillators (such as high energy physics, positron emission tomography, and  $\gamma$ -ray detection), a high light yield is important along with good energy resolution, and fast response time. However, the search for new scintillators has now slowed down. Therefore, the attention again turned to materials that were previously considered to be lumino-phores. Moreover, it appears that scintillators based on binary systems have a higher light output than the original matrices. The light output for BaBrI:Eu<sup>2+</sup> crystals was estimated as 90000 photons/MeV [1], whereas in BaBr<sub>2</sub> and BaI<sub>2</sub> it is only 58000 and 22000 photons/MeV, respectively [2]. Light output of SrBrI single crystal doped with Eu<sup>2+</sup> was estimated about 47000 photons/MeV [1]. However, optical properties of these crystals has not been investigated well. In a previous article we have studied optical properties

of BaBrI and BaClI crystals. The exciton creation energy, band gap and the energy of the lowest 4f-5d transition were estimated and VRBE diagram were constructed [3].

In this article, we present study of the optical properties of pure and Eu-doped SrBrI as well as BaBrI crystals. The optical properties of oxygen centers in BaBrI and SrBrI crystals are also reported. The band gap of SrBrI crystal is estimated. Thermal quenching of Eu<sup>2+</sup> luminescence is studied. The vacuum referred binding energy (VRBE) diagram for SrBrI crystal is constructed based on experimental data and GGA-PBE calculations.

## 2. Methodology

### 2.1. Crystal growth

The compounds of alkaline earth halides are hygroscopic. Therefore, much attention is paid to drying raw materials before crystal growth. The thermogravimetric method (TG) and differential scanning calorimetry (DSC) were used to determine the melting points, the level of hydration and the possible dehydration temperature of the charge materials. The analysis of the charge was carried out using a synchronous thermal analyzer STA 449 Jupiter (NETZSCH). The starting materials were the anhydrous compounds

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of BaBr<sub>2</sub>, BaI<sub>2</sub>, SrBr<sub>2</sub> and SrI<sub>2</sub> (purity 99.9%, Lanhit, LTD). The stoichiometric mixtures of BaBr<sub>2</sub>+BaI<sub>2</sub> and SrBr<sub>2</sub>+SrI<sub>2</sub> were employed. The EuBr<sub>3</sub> was used for doping the mixtures. Eu<sup>2+</sup> concentration employed is 0.05 mol.%. The melting points for BaBrI and SrBrI are 783 °C and 507 °C, respectively.

Quartz ampoules with special design of 20–30 mm in diameter were developed for bromide mixture drying and melt filtration. This procedure significantly improves the optical quality and scintillation characteristics of the crystals. Quartz ampoule consisted of two sections separated by quartz filter.

The thermal unit with controlled heating was used to dry the batch in the first section of quartz tube under vacuum. The drying process for raw materials was performed taking into account the TG and DSC data. As a next step, the charge was melted after thorough drying.

The molten compounds were filtered through a quartz filter. The filtered melt was collected in the second section of the ampoule. This section was sealed off under vacuum after filtration and cooling of the melt. The synthesized compound into this quartz ampoule was used to grow a single crystal. A vertical 20-zone furnace was used to grow BaBrI:Eu and SrBrI:Eu single crystals via the Bridgman method. The temperature gradient was 10–15 °C/cm. The rate of the crystal growth was 1 mm/h. The cooling rate of grown crystals was 10 °C/h. After that the growth procedure, the crystals were stored in mineral oil to protect from the atmosphere contamination.

Irregularly shaped 1–1.5 mm thick and 10–15 mm diameter crystal pieces were cleaved from the original boules for further studies. The plates were polished in glove box for the spectroscopic measurements.

## 2.2. Luminescence measurements

Photoluminescence (PL) was measured in vacuum cold-finger cryostat. The spectra were recorded with a MDR2 and SDL1 (LOMO) grating monochromator, a photomodule Hamamatsu H6780-04 (185–850 nm), and a photon-counter unit. The luminescence spectra were corrected for spectral response of detection channel. The photoluminescence excitation (PLE) spectra were measured with a grating monochromators MDR2 and 200 W xenon arc lamp for direct 4f-5d excitation and vacuum monochromator VM-2 (LOMO) and Hamamatsu deuterium lamp L7292 for measurements in VUV spectral region. The PLE spectra were corrected for the varying intensity of exciting light due. The X-ray excited luminescence was performed using an X-ray tube operating at 50 kV and 1 mA. Temperature dependences of luminescence were recorded in vacuum cryostat at linear heating. The heating rate was about 10 K/min.

## 2.3. Calculation details

*Ab Initio* calculations of BaBrI and SrBrI doped with Eu<sup>2+</sup> were carried out within density functional theory (DFT) using VASP (Vienna *Ab Initio* Simulation Package) code [4]. The 2 × 2 × 1 (48 atoms) supercell was constructed, in which one of Ba<sup>2+</sup> or Sr<sup>2+</sup> ions was replaced by Eu<sup>2+</sup>.

Integration within the Brillouin zone was performed on a  $\Gamma$ -centered grid of 8 irreducible  $k$  points. Geometry optimization was performed with fixed cell dimensions. The convergence was achieved if the difference in total energy between the two last iterations was less than 10<sup>-6</sup> eV.

## 3. Results and discussion

### 3.1. Exciton emission

The wide emission band at about 3.7 eV was observed at 78 K in undoped SrBrI crystals when excitation was performed into 5–10 eV region. The excitation peak (Fig. 1, curve 1) was located at about 5.6 eV. At 5.7 eV a small dip was monitored in excitation spectrum. Intensity of this band decreased in the doped crystals.

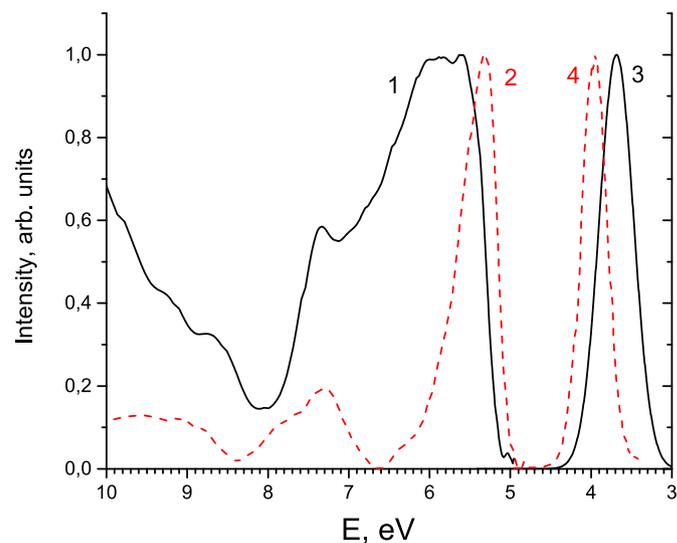
In BaBrI crystals the wide band luminescence peaked at 3.9 eV was observed. The excitation peak is located at about 5.3 eV. The wide peaks were found in excitation spectra in 5.5–10 eV range. This luminescence were attributed to self-trapped exciton luminescence [3]. The same luminescence were observed in the crystals under x-ray excitation.

In SrBrI the most efficient photoluminescence is excited within interband absorption, but excitation peak is located in a higher energy in comparison with BaBrI whereas emission band is shifted to lower energy. This behavior is typical for self-trapped exciton (STE). Similar to other alkali-earth halides [5–7], the STE in SrBrI can consist of molecular ion similar to H-center (hole on interstitial bromine or iodine ion) and an F-center-like part (electron trapped by iodine or bromine vacancy). Due to presence of different non-equivalent positions it is possible to observe several types of STE. However, in the present article no attention was paid to this question.

Early we estimated band gap of BaBrI and SrBrI crystals using position of exciton peak. Energy of exciton peak in SrBrI is about 5.7 eV. Assuming binding energies of excitons in SrBrI and BaBrI are close, we can estimate band gap of SrBrI crystal following the procedure proposed in Ref. [3]. Exciton creation energy ( $E_x$ ) in SrBrI is obtained at about 5.7 eV. Therefore, band gap is estimated about 6.0 eV.

### 3.2. Eu<sup>2+</sup>-luminescence

The Eu doping leads to appearance of intense luminescence band peaked at about 2.94 eV. This emission spectrum results from 5d-4f transitions in Eu<sup>2+</sup> ions. The emission peak is slightly shifted



**Fig. 1.** Excitation and emission spectra of self-trapped exciton in BaBrI and SrBrI at 78 K. Excitation spectra of SrBrI (solid, curve 1) and BaBrI (dashed, curve 2) crystals monitored at 3.65 eV and 3.95 eV, respectively. Emission spectra were excited at 5.6 eV in SrBrI (curve 3) and 5.3 eV in BaBrI (curve 4) crystals.

to the lower energy region in comparison with  $\text{Eu}^{2+}$  emission band in BaBrI crystals at 3.0 eV. Excitation spectra of the emission of Eu-doped BaBrI and SrBrI crystals are given in Fig. 2. The intense excitation peak at 4 eV is attributed to 4f-5d transition in  $\text{Eu}^{2+}$  ions in SrBrI crystals. The energy of 4f-5d transition in SrBrI is slightly lower than in BaBrI crystals. Using method proposed by Dorenbos [8] the lowest energy of 4f-5d transition in BaBrI was estimated as 3.29 eV [3]. The excitation peak of 5d-4f luminescence in SrBrI shifts by 0.2 eV to the lower energies with respect to BaBrI. Therefore, the lowest energy of 4f-5d transition can be estimated about 3.1 eV.

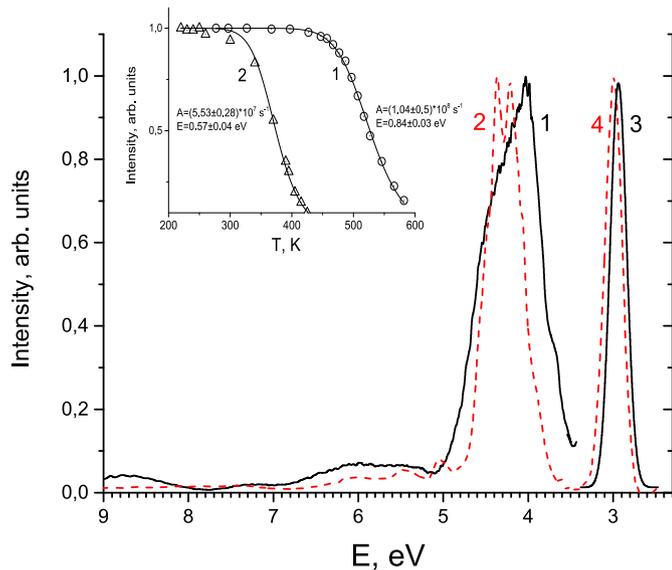
Besides the intense 4f-5d excitation band we found that 5d-4f  $\text{Eu}^{2+}$  emission was excited in exciton and interband transitions region (see Fig. 2). This provides the evidence of energy transfer from the excitons and primary energy carriers (electrons and holes) to the  $\text{Eu}^{2+}$  ions.  $\text{Eu}^{2+}$  emission is more intensive in SrBrI than in BaBrI under exciton excitation. It is due to more overlapping of 4f-5d absorption band and exciton emission in SrBrI than in BaBrI crystals.

The intensity of  $\text{Eu}^{2+}$  luminescence under intra center direct 4f-5d excitation decreases with increasing temperature. The temperature dependences of  $\text{Eu}^{2+}$  luminescence for SrBrI and BaBrI crystals are given in the inset in Fig. 2. In BaBrI crystals  $\text{Eu}^{2+}$  emission is slightly quenched already at room temperature and intensity dropped to 50% of the low temperature value at 380 K ( $T_{0.5} = 380$  K). In SrBrI crystals the  $\text{Eu}^{2+}$  emission is quenched at higher temperatures and  $T_{0.5}$  value is about 520 K. The luminescence is quenched following Mott's law:

$$I(T) = \frac{1}{1 + w_0 \exp(-E_A/k_B T)}, \quad (1)$$

where  $w_0$  is the rate constant for the thermally activated escape and  $E_A$  is the activation energy connected with this process. The energies are  $0.84 \pm 0.03$  eV and  $0.5 \pm 0.04$  eV for SrBrI and BaBrI, respectively.

The different mechanisms of  $\text{Eu}^{2+}$  luminescence quenching



**Fig. 2.** Excitation and emission spectra of  $\text{Eu}^{2+}$  doped BaBrI and SrBrI at 78 K. Excitation spectra of SrBrI (solid, curve 1) and BaBrI (dashed, curve 2) crystals monitored at 3.0 eV and 2.9 eV, respectively. Emission spectra were excited at 4.0 eV in SrBrI (curve 3) and 4.2 eV in BaBrI (curve 4) crystals. In the inset temperature dependences of 5d-4f luminescence intensity of  $\text{Eu}^{2+}$  ions under 4.0 eV excitation in SrBrI (curve 1) and BaBrI (curve 2) crystals are given.

exist. One of the first models proposed by Blase pointed that the quenching of 5d-4f emission in  $\text{Eu}^{2+}$  could be attributed to a large displacement between the ground and excited states of  $\text{Eu}^{2+}$  in the configuration coordinate diagram [9]. In the other model Davolos [10] assumed that the thermal activation from the  $4f^6 5d^1$  state to conduction band state causes the quenching in Ba thiogallates. Based on an analysis of the temperature quenching of  $\text{Eu}^{2+}$  emission in numerous of compounds, Dorenbos showed that the quenching is not due to a large displacement between the ground and excited states of  $\text{Eu}^{2+}$  in the configuration coordinate diagram. It was proposed that thermal excitation of the 5d electron to conduction band states appears the genuine mechanism. The 5d electron remains bonded in an  $\text{Eu}^{3+}$  trapped exciton state from which it returns non-radiative to the  $\text{Eu}^{2+}$  ground state [11].

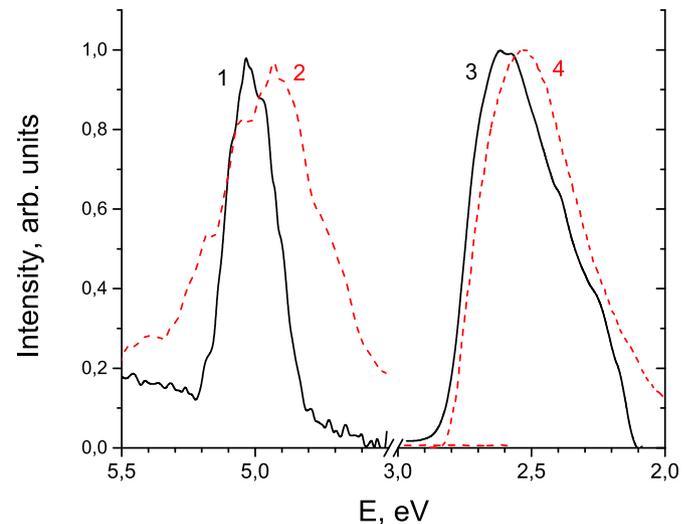
### 3.3. Luminescence related to oxygen centers

In some nominally undoped samples of SrBrI and BaBrI crystals the wide band emission peaked at about 2.5 eV was observed. The position of emission and excitation bands slightly depends on the cation type. In SrBrI crystal the wide emission band peaked at 2.6 eV was excited in the band centered at 5.0 eV. In BaBrI crystals emission band centered at 2.5 eV. In the excitation spectrum monitored at 2.5 eV the band peaked at 4.9 eV was found (Fig. 3).

The intensity of this emission band depends on preparation of charge for crystal growth. An intensity of this luminescence band was higher in the samples grown by the fast dried and unfiltered raw, where insufficient dehydration took place. When  $\text{H}_2\text{O}$  molecules were removed completely from the raw after melt filtration and long 24-h drying procedures, the intensity of the 2.5 eV emission band is quite low.

This emission bands in BaClI, BaBrI and SrBrI crystals are located in the same spectral region and can be attributed to the oxygen centers [3]. In strontium iodide a significant part of the oxygen contamination comes from the surface hydrate reactions [12]. Therefore, concentration of oxygen centers in crystals grown from insufficiently dried substrate should be higher.

Several types of oxygen centers were found in alkali halide and alkali-earth crystals. In alkali-earth fluorides ( $\text{CaF}_2$ ,  $\text{SrF}_2$ ,  $\text{BaF}_2$ ) and alkali earth dihalides BaFCl, BaFBr oxygen centers ( $\text{O}_X^{2-} - \nu_X^-$ , where



**Fig. 3.** Excitation and emission spectra of oxygen centers in nominally undoped SrBrI and BaBrI crystals. Excitation spectra of SrBrI (solid, curve 1) and BaBrI (dashed, curve 2) crystals monitored at 2.6 eV and 2.5 eV, respectively. Emission spectra were excited at 5.0 eV in SrBrI (curve 3) and 4.9 eV in BaBrI (curve 4) crystals.

X = F, Cl, Br) are formed [13–17]. Oxygen ion replaces fluorine ion. Charge compensation is performed via anion vacancy formation. The  $O_X^{2-} - v_X^{\cdot}$  centers have several absorption bands and one or several broad luminescence bands. It is established that the lowest-energy excitation (absorption) band is approximately constant in crystals with the same alkali metal and follows the Molvo-Iv dependence in crystals with the same halogen [16,18].

In the most of alkali-halide crystals, molecular ions of oxygen ( $O^{2-}$ ) were found. The oxygen ions have several absorption bands at about 4–6 eV and a number of well-resolved emission lines at about 1–2 eV [19]. The strong ESR signal should be observed in crystals containing oxygen in molecular ion form. The entry form of oxygen is host material related. For example, molecular oxygen ions are practically not observed in alkaline-earth crystals, whereas centers containing oxygen ion and anion vacancy are not observed in alkali-halide crystals [13,16].

Oxygen centers related luminescence was observed in BaBrF and BaClF crystals. The oxygen ion ( $O^{2-}$ ) can occupy  $Br^-$  or  $F^-$  position in these crystals. Therefore, several types of the centers could be found, there were  $O_X^{2-} - v_X^{\cdot}$ ;  $O_X^{2-} - v_X^{\cdot}$  (type II centers);  $O_F^{2-} - v_X^{\cdot}$  (type I centers);  $O_F^{2-} - v_X^{\cdot}$  (type III centers), where X = Br, Cl. Here  $v_X^{\cdot}$  is vacancy of the anion X. Type I centers, where oxygen substitutes for fluorine ion and anion vacancy occupies bromine or chlorine site, have low energy absorption band at 5.0 eV and wide luminescence band peaked at 2.5 eV [15]. Type II centers, where oxygen ion and anion vacancy are located in bromine sublattice, demonstrate absorption band at about 4.2 eV and luminescence at about 2 eV [14]. Type III centers were not clearly separated in optical absorption spectra and detected only in EPR [20]. Ground state of type I centers is mostly formed by 2p orbitals of oxygen and is located in the gap between fluorine and bromine subbands that formed valence band. The observed luminescence peaked at 2.5 eV and excitation band are very similar to luminescence of type I centers in BaFBr crystals. Intensity of the luminescence depends on dehydration process. Therefore, we can conclude that the 2.5 eV luminescence is related to oxygen vacancy centers. Following to the BaFBr crystals we can expect, that ground state consists from the 2p orbitals of oxygen ions substituting iodine ions but first excited state can be attributed to 1s state of anion vacancy occupied bromine site. The energy of first excitation band is close to band gap. Therefore, if the ground state of oxygen center was in band gap the luminescence would be quenched at room temperature. However, this does not happen in our crystals. Therefore, the ground state of centers containing oxygen ion and anion vacancy should be placed into the gap between iodine and bromine subbands.

#### 3.4. Calculation results and vacuum referred binding energy (VRBE) diagram of the lanthanide ions in SrBrI

The ground state of crystals was calculated using spin-polarized generalized gradient approximation (GGA) with the exchange-correlation potential PBE [21]. The electronic structure parameters of GGA-PBE calculations are often underestimated, which turned out to be in our case. The calculated band-gap values of the

crystals are presented in Table 1 together with experimental data.

For the calculation of doped crystals, we replaced one of  $Ba^{2+}$  ions in the supercell by the europium ion. To estimate the gap between top of valence band and ground 4f state of  $Eu^{2+}$  ion ( $E_{CT}$ ), we applied the method described in Ref. [22] using the Dudarev's approximation PBE + U [23] for adjusting deposition of 4f levels. In accordance with our previous calculations [3], we found that the adequate value of  $U_{eff}$  is 2.5. The results of PBE + U calculations are presented in Table 1. The gap between lowest 5d state of  $Eu^{2+}$  and bottom of conduction band ( $E_{dC}$ ) were performed using following relation:

$$E_{dC} = E_g - E_{CT} - E_{fd}, \quad (2)$$

where  $E_{dC}$  – energy gap between the lowest  $4f^6 5d^1$  state and bottom of conduction band,  $E_{CT}$  – calculated energy of  $Eu^{3+}$  charge transfer band, and  $E_{fd}$  – experimental energy of the first f-d transition in  $Eu^{2+}$  ions.

Despite we used *ab initio* methodology, our calculations are to some extent empirical. So we reexamine the calculation using the GW0 [24] and BSE (Bethe-Salpeter equation) [25] methods. The values of the band gap calculated using the GW0 method gives good agreement with the experiment. From the spectra of the imaginary part of dielectric function calculated by BSE method, we estimated the 4f-5d transition energies. All calculated data are presented in Table 1. Both methods of calculation give  $E_{dC}$  energies of about 0.8 eV, which indicates the promise of SrBrI crystals as scintillators. The calculated  $E_{dC}$  energies are in good agreement with the experimental ones estimated in the inset of Fig. 2. The calculated values of band gap energies agree well with the experimental data (Table 1). Therefore, the performed calculation could be used for estimation of band gap and position of 4f levels in other new alkali-earth halides.

Using measured energies of band gap, exciton creation and 4f-5d transition we can construct vacuum referred binding energy diagram. The diagram shows the binding energies of an electron in the divalent and trivalent lanthanide ion ground and excited states using Dorenbos chemical shift model [26]. We have no spectroscopic data on trivalent lanthanides in investigated crystals, therefore we use a Coulomb repulsion energy  $U(6, A)$  which defines the binding energy difference between an electron in the ground state of  $Eu^{2+}$  with that in the ground state of  $Eu^{3+}$  [27,28] to define electron binding energy of trivalent lanthanides:

$$U(6, A) = E_{4f}(7, 2+, A) - E_{4f}(6, 3+, A). \quad (3)$$

The ground levels of the 5d configuration of trivalent lanthanides were estimated from the relation within the redshift models [8,27]:

$$D(3+, A) = 1.563 \cdot D(2+, A) + 0.364, \quad (4)$$

where  $D(3+, A)$  is redshift of trivalent lanthanide and  $D(2+, A)$  is redshift of divalent lanthanide. Similar to other alkali-earth iodide and bromide compounds like  $LaBr_3$   $U(6, A)$  is estimated about 6.3 eV. The positions of 4f and 5d levels of trivalent lanthanides could be estimated using Eq.(3) and (4).

Dorenbos showed [11] that the energy barrier of thermal quenching ( $E_A$  from Eq. (1)) of 5d-4f emission is related to energy gap between the lowest  $4f^6 5d^1$   $Eu^{2+}$  state and the bottom of conduction band. In SrBrI and BaBrI the energy barriers for the luminescence quenching ( $E_A$ ) agree well with the calculated values ( $E_{dC}$ , see Eq. (2)) as shown in Table 1. Calculated energy  $E_{CT}$  for  $Eu^{2+}$  ion in BaBrI and SrBrI has lower energy than 4f-4f emission of  $Eu^{3+}$  ions which means the absence of charge transfer transition in  $Eu^{3+}$

**Table 1**

Calculated band gaps and relative 4f and 5d levels for  $Eu^{2+}$ -doped BaBrI and SrBrI crystals. Energies are given in eV.

Crystal	Band gap		$E_{CT}$	$E_{fd}$	$E_{dC}$ calc.	$E_A$ exp.
	GW <sub>0</sub>	Exp.				
BaBrI	5.34	5.58	1.4	3.29	~ 0.65	0.57 ± 0.04
SrBrI	5.5	6.00	1.62	3.1	~ 0.85	0.84 ± 0.03

excitation spectrum in these crystals. Therefore, the calculated charge transfer energy  $E_{CT}$  and experimental values  $E_A$  can be used to estimate the energy gap between the lowest  $4f^65d^1$  state and bottom of conduction band and position of top of valence band related to  $\text{Eu}^{2+}$  4f ground state.

The VRBE diagram for SrBr crystals are given in Fig. 4. The density of states is calculated. It is found the small gap between bromine and iodine subzone similar to BaFBr, BaFCl crystals [15,20]. Following the diagram we can expect the bright luminescence 5d–4f from  $\text{Sm}^{2+}$  and  $\text{Yb}^{2+}$  ions in SrBr crystals. It would be promising for creation the fast scintillators [29,30]. The Ce-doping is also prospecting way due to 5d state of  $\text{Ce}^{3+}$  ion could be located in sufficiently far from the bottom of conduction band to avoid quenching. We estimate that the most ground 4f states of trivalent lanthanide ions except  $\text{Ce}^{3+}$  lie deep in the valence band. Therefore, we cannot expect to find 5d–4f luminescence of these lanthanides.

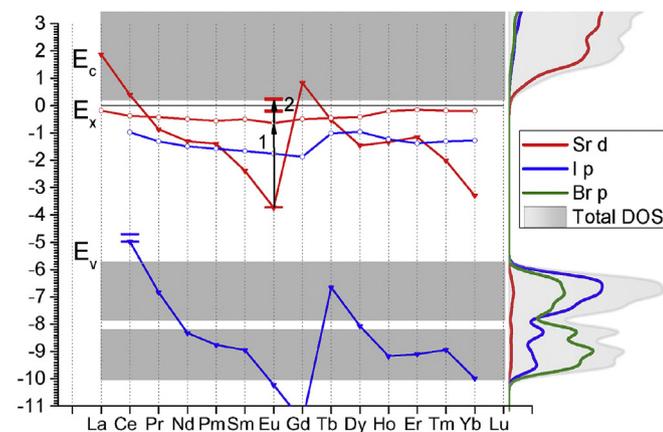
#### 4. Conclusion

The SrBr and BaBr single crystal undoped and doped with 0.05 mol.%  $\text{Eu}^{2+}$  ions were grown. The luminescence properties of SrBr and BaBr crystals doped with low concentrations of  $\text{Eu}^{2+}$  ions have been studied. The results obtained by luminescence spectroscopy revealed the presence of europium ions only in divalent form in the BaBr and SrBr crystals.

The intense bands in excitation spectra of SrBr–Eu crystals at about 3.85–4.2 eV result from  $4f^7(^8S_{7/2}) \rightarrow 4f^65d^1(t_{2g})$  transitions. The lowest energies of 4f–5d transition in  $\text{Eu}^{2+}$  ion obtained from the spectra was 3.1 eV for SrBr. The narrow intense peak at 5.7 eV in the photoluminescence excitation spectrum of the intrinsic emission of SrBr crystal was due to the creation of excitons. The band gap of SrBr crystal were estimated about 6.0 eV.

In the undoped crystals wide band luminescence at about 2.5 eV was observed under 5 eV excitation. It corresponds to the centers containing oxygen ion and anion vacancy.

The band gap was estimated using the  $\text{GW}_0$  approximations. The calculated band gap energies agree with the experimental data.



**Fig. 4.** The diagram shows the vacuum referred binding energy (VRBE) for electrons of divalent (red) and trivalent (blue) lanthanide ions with respect to the vacuum level in SrBr. The zigzag red curve with triangles connects the  $4f^n$  ground state energies of the divalent lanthanide ions. The red curve with circles is attributed to their lowest  $4f^{n-1}5d^1$  states. Blue zigzag curve with triangles respects to  $4f^n$  ground state energies of trivalent lanthanide ions and blue curve with circles are their  $4f^{n-1}5d^1$  states. Arrow 1 shows the energy of 4f–5d transition and arrow 2 shows the gap between the lowest  $4f^65d^1$  state and bottom of conduction band estimated using experimental  $E_A$  (Eq. (1)) and calculated  $E_{AC}$  (Eq. (2)) energies. The right inset provides the calculated density of states (DOS) for SrBr crystal.

The distance between the lowest  $4f^65d^1$  level of  $\text{Eu}^{2+}$  ions and top of the valence band has been calculated and estimated using energy barrier of 5d–4f  $\text{Eu}^{2+}$  luminescence quenching. The calculated and experimental values of energy gap between the lowest  $4f^65d^1$  state and bottom of conduction band are in good agreement. The VRBE diagrams of levels of all divalent and trivalent lanthanide ions in SrBr crystals were constructed on the acquired experimental and theoretical data.

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